

**YOU MAY USE A SEPARATE PIECE OF PAPER TO COMPLETE THIS**

REMINDERS:

- Metals always carry a positive charge
- Nonmetals have negative charges when they are in an ionic compound
- Silver is always a +1 charge
- Cadmium and Zinc are always a +2 charge

IONIC BONDS:

- Transfer valence electrons
- Metals and non metals
- metals and polyatomics
- polyatomics & nonmetals
- polyatomics and polyatomics

COVALENT BONDS:

- Share electrons valence
- Between two different nonmetals OR
- Between two of the same nonmetal (called “diatomic molecules”)

HYDROGEN – usually carries a positive charge, but it is NOT a metal

**Write the appropriate chemical name for the compounds: Place correct charges for Ionic Compounds only!**

1. NaCl
2. KBr
3. CaF<sub>2</sub>
4. FeO
5. Fe<sub>2</sub>O<sub>3</sub>
6. CuCl<sub>2</sub>
7. NO<sub>2</sub>
8. CO
9. C<sub>2</sub>H<sub>6</sub>
10. Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>
11. NaOH
12. FeCrO<sub>4</sub>
13. Pb<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>
14. P<sub>2</sub>O<sub>5</sub>
15. Ca(ClO<sub>3</sub>)<sub>2</sub>
16. (NH<sub>4</sub>)<sub>2</sub>O
17. Zn(HCO<sub>3</sub>)<sub>2</sub>
18. SnBr<sub>4</sub>
19. (NH<sub>4</sub>)<sub>2</sub>S
20. NiI<sub>2</sub>

**Write the appropriate chemical formula for the following compounds: Place correct charges in the formula for Ionic Compounds only!**

21. Lithium oxide
22. Aluminum sulfide
23. Calcium chloride
24. Lead (IV) oxide
25. Copper (II) iodide
26. Mercury (II) hydroxide
27. Dinitrogen pentoxide
28. Carbon tetrahydride
29. Dihydrogen monoxide
30. Ammonium chloride
31. Copper (I) sulphate
32. Sodium phosphate
33. nickel(II) nitrate
34. lithium chromate
35. potassium permanganate
36. silver perchlorate
37. potassium phosphate
38. sodium nitrate
39. silver sulfide
40. nickel(II) acetate